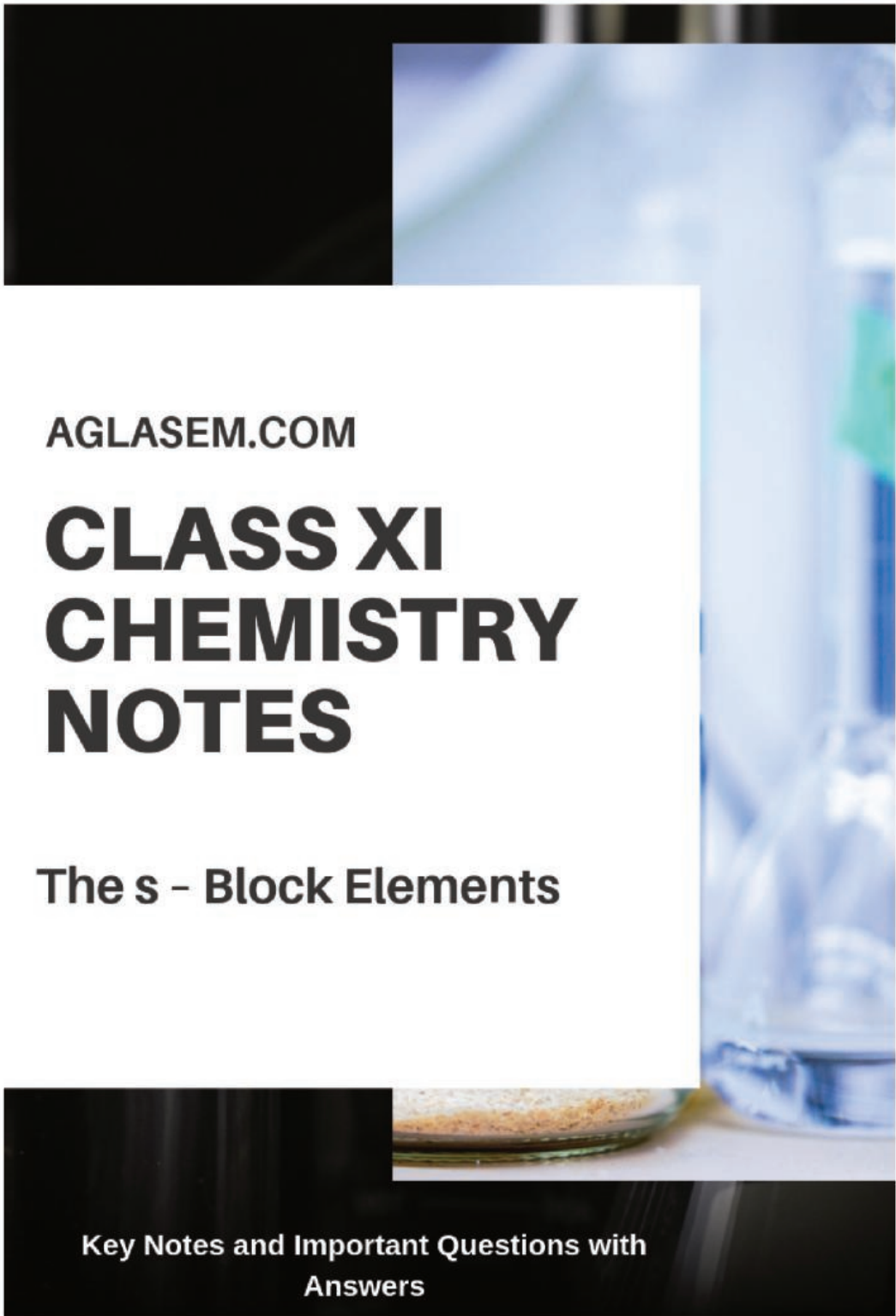


The background of the cover features a photograph of laboratory glassware. On the right side, there is a tall, clear glass bottle or flask. In the foreground, at the bottom, there is a smaller glass dish or beaker containing a yellowish-orange substance. The image is partially obscured by a large white rectangular area on the left where the text is located.

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CLASS XI CHEMISTRY NOTES

The s - Block Elements

**Key Notes and Important Questions with
Answers**



Chapter - 10

s-Block Elements

- s-block elements consists of group-I (Alkali metals) and group-2 (Alkaline earth metals).
- Group 1st elements—Li, Na, K, Rb, Cs, Fr.
- Group 2nd elements—Be, Mg, Ca, Sr, Ba, Ra.
- **Atomic radius**—Atomic radius of alkali metals are greater than alkaline earth metals.
- **Hydration enthalpy** : Decreases with increases in ionic sizes.
- **Ionic mobility** : Smaller the size of ion, more highly it is hydrated and hence lower is its ionic mobility.



- **Ionisation enthalpies** : 1st I.E. of group 1st is smaller than group 2nd elements but 2nd I.E. of group 2nd is smaller than group 1st elements.
- **Flame colouration** : Due to low I.E., s-block elements and their salts imparts characteristics colour of oxidising flame (except Be and Mg). Be and Mg do not show flame colouration because they have small size and very high ionisation enthalpy.
- **Reducing character** : Due to large negative electrode potentials alkali metals are stronger reducing agent than alkaline earth metal.

- **Reactivity towards air** :



Alkaline earth metals being smaller in size do not form superoxides.

- **Reactivity towards H₂O** :

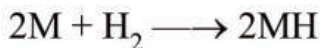


(Alkali metal)



(Alkaline earth metals)

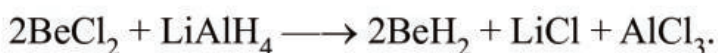
- **Reactivity towards hydrogen :**



(M = Li, Na, K, Rb, Cs)



(M = Mg, Ca, Sr, Ba)



- **Reactivity towards halogens :**



- **Solution in liquid ammonia :** The fresh solution of alkali metals and alkaline earth metals (except Be and Mg) is deep blue, paramagnetic and highly reducing due to presence of ammoniated electrons.

- **Solubility of alkaline earth metal carbonate in water :**



- **Solubility of alkaline earth metal carbonates in water.**



- **Solubility of alkaline earth metal sulphates in water :**



- **Thermal stability of alkali metal carbonates :**



- **Thermal stability of alkaline earth metal carbonates :**



- **Anomalous behaviour of Li and Be :** It is due to very small size, high I.E. and high polarising power (*i.e.*, charge/radius)

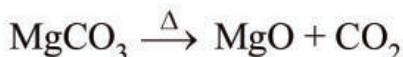
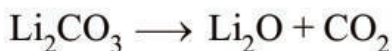
- **Diagonal relationship (similarities) between Li and Mg :**

(i) Both Li and Mg are hard.

(ii) Both react with N_2 to form nitrides.



(iii) **Decomposition of carbonates :**



(iv) Both LiCl and MgCl_2 are deliquescent. They form hydrate salts $\text{LiCl} \cdot 2\text{H}_2\text{O}$ and $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$.

(v) **Decomposition of nitrates :**



● **Diagonal relationship (similarities) between Be and Al :**

(i) Both are passive to acids due to formation of oxide layer.

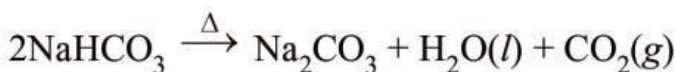
(ii) Hydroxides of both dissolve in alkali to form $[\text{Be}(\text{OH})_4]^{2-}$ and $[\text{Al}(\text{OH})_4]^-$.

(iii) Chloride of both has bridged structure.

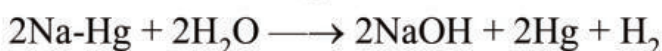
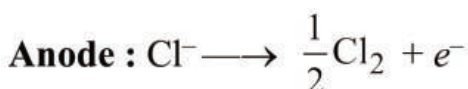
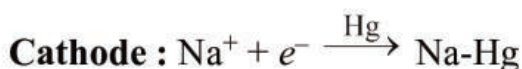
(iv) Both have tendency to form complexes of BeF_4^{2-} , AlF_6^{3-} .

● **Manufacturing of washing soda ($\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$) :**

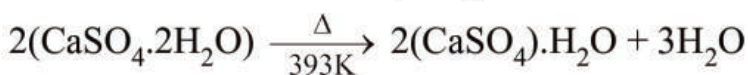
Solvay process :



● **Manufacturing of caustic soda (NaOH) :** Castner-Kellner cell.



● **Plaster of paris :** ($\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$)



Gypsum

- **Cement** is a finely powdered mixture of calcium silicates and aluminate along with small quantities of gypsum which sets into a hard stone like mass when treated with water.

1 - Mark Questions

1. What is the oxidation state of K in KO_2 ?
2. Why are group I element called alkali metals ?
3. Potassium carbonate cannot be prepared by solvay process. Why ?
4. LiCl is soluble in organic solvent. Why ?
5. Why are group I elements called alkali metals ?
6. Alkali metals are strong reducing agents. Why ?
7. Why do alkali metals give characteristics flame colouration ?
8. Arrange the following in order of increasing covalent character : MCl , MBr , MF , MI (where M = Alkali metal) [**Ans.** $\text{MF} < \text{MCl} < \text{MBr} < \text{MI}$]
9. Alkali metals can not be obtained by chemical reduction method. Explain.
10. Why is sodium metal kept under kerosene oil ?
11. Why Be and Mg do not give characteristics colour to the flame ?
12. Arrange the alkaline earth metal carbonate in the decreasing order of thermal stability.
13. Why do alkaline earth metals not form any superoxide ?
14. Why gypsum is added to cement ?
15. How plaster of paris is obtained from gypsum ?
16. BeO is insoluble in water but BeSO_4 is soluble in water ? Why ?
17. Why second I.E. of group II elements is less than group I elements ?
18. What is quick lime ? How is it prepared ?
19. Why does Be shows similarities with Al ?
20. Name the alkaline earth metal hydroxide which is amphoteric.

2 - Mark Questions

1. Why are alkali metals soft and have low melting points ?
2. Write any four similarities between Li and Mg.
3. Why are potassium and caesium rather than Lithium used in photoelectric cells ?
4. Why is Li_2CO_3 decomposed at a lower temperature whereas Na_2CO_3 at higher temperature ?
5. Among the alkali metals which has :
 - (i) Highest melting point.
 - (ii) Most electropositive character
 - (iii) Lowest size of ion.
 - (v) Strongest reducing character.[Ans. (i) Li (ii) Cs (iii) Li (iv) Li]
6. Why does the solubility of alkali earth metal carbonates and sulphates decreases down the group ?
7. Draw the structure of BeCl_2 in (i) Vapour phase (ii) Solid state.
8. When CO_2 gas is passed in lime water it turns milky but in case of excess CO_2 milkiness disappears. Support the statement by giving suitable reaction equations.
9. (i) E^θ for $\text{M}^{2+}(\text{aq}) + 2e^- \longrightarrow \text{M}(\text{s})$ (where $\text{M} = \text{Ca}, \text{Sr}, \text{Ba}$) is nearly constant.
(ii) What is dead burnt plaster ? How is it obtained from gypsum?
10. Write two important uses of (i) Limestone (ii) Quick lime.

3 - Mark Questions

1. Assign reason for the following :
 - (i) Compounds of lithium are generally covalent.
 - (ii) Alkali metals are strong reducing agent.
 - (iii) LiCl is more covalent than NaCl .
2. Discuss the various reactions that occur in Solvay process.

3. Explain why ?
- (i) Lithium salts are commonly hydrated.
 - (ii) Sodium peroxide is widely used as an oxidising agent.
 - (iii) Sodium wire is used to remove moisture from benzene but can't be used for drying alcohol.
4. Sodium hydroxide is generally prepared by electrolysis of brine solution in the Castner-Kellner cell :
- (i) Write the reactions that occur in the cell.
 - (ii) Write any two uses of NaOH.
5. Explain with suitable reasons :
- (a) A solution of Na_2CO_3 is alkaline.
 - (b) Alkali metals are prepared by electrolysis of their fused chlorides.
 - (c) Sodium is found to be more useful than potassium ?
6. Arrange the following in order of property mentioned against each :
- (i) BaCl_2 , MgCl_2 , BeCl_2 , CaCl_2 (Increasing ionic character)
 - (ii) $\text{Mg}(\text{OH})_2$, $\text{Sr}(\text{OH})_2$, $\text{Ba}(\text{OH})_2$, $\text{Ca}(\text{OH})_2$ (Increasing solubility in water)
 - (iii) BeO , MgO , BaO , CaO (Increasing basic strength)
7. What happens when :
- (i) Mg is burnt in air.
 - (ii) Quick lime is heated with silica.
 - (iii) Chlorine is heated with slaked lime.
8. Write the raw material required for the manufacture of portland cement ?
Why gypsum is added into it.
9. (i) Why alkaline earth metals cannot be obtained by reduction of their oxide ?
- (ii) Why the elements of group 2 are known as alkaline earth metals ?
10. (i) Alkaline earth metals form ionic salt having bivalent cations. Explain. Why ?
- (ii) A piece of magnesium ribbon continues to burn in SO_2 . Why ?

5 - Mark Questions

1. Explain the following observation :

- (a) LiI is more soluble than KI in ethanol.
- (b) Sodium reacts with water less vigorously than potassium.
- (c) LiF is insoluble in water.
- (d) The mobilities of the alkali metal ions in aqueous solution are $\text{Li}^+ < \text{Na}^+ < \text{K}^+ < \text{Rb}^+ < \text{Cs}^+$.
- (e) Lithium is the only alkali metal to form a nitride directly.

2. Complete the following reaction equations :

- (i) $\text{BeCl}_2 + \text{LiAlH}_4 \rightarrow$
- (ii) $\text{CaO} + \text{SiO}_2 \rightarrow$
- (iii) $\text{Ca(OH)}_2 + \text{Cl}_2 \rightarrow$
- (iv) $\text{CaO} + \text{P}_4\text{O}_{10} \rightarrow$
- (v) $\text{Ca(OH)}_2 + \text{CO}_2 \rightarrow$

3. Compare the solubility and thermal stability of the following :

Compounds of the alkali metals with those of alkaline earth metals
(a) nitrates (b) carbonates (c) sulphates.

4. Explain the significance of Sodium (Na), Potassium (K), Magnesium (Mg) and Calcium (Ca) in biological fluids.

5. Explain the significance of Sodium Potassium, Magnesium and Calcium biological fluids.

6. (i) A solutions of Na_2CO_3 is alkaline why?

(ii) BeO insoluble but BeSO_4 in soluble in water. Why?

(iii) Lithium salts are commonly hydrated and those of other alkali metal ions are usually anhydrous give reasons.

(iv) What is the importance of cement.

(v) What happen when quick lime is heated with silica.

