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CLASS XII CHEMISTRY NOTES

CHEMICAL KINETICS

KEY NOTES AND IMPORTANT
QUESTIONS WITH ANSWERS



UNIT 4

CHEMICAL KINETICS

CONCEPT

Chemical Kinetics : The branch of physical chemistry which deals with the study of rate of reaction and factors affecting rate.

Rate of chemical reaction : The change in concentration of any reactant or product per unit time is called rate of reaction.

Types of Rate of Reaction :

1. **Average rate of reaction :** The rate of reaction measured over the long time interval is called average rate of reaction.

$$\text{Avg rate } \frac{\Delta x}{\Delta t} = \frac{-\Delta[R]}{\Delta t} = \frac{\Delta[P]}{\Delta t}$$

2. **Instantaneous rate of reaction :** The rate of reaction measured at a particular time is called instantaneous rate of reaction.

$$(\text{Rate})_i = (\text{Instantaneous rate}) \, dx/dt = -d[R]/dt = +d[P]/dt$$

Factors affecting Rate of Reaction

1. Concentration of reactant
2. Surface area
3. Temperature
4. Nature of reactant
5. Presence of catalyst
6. Radiation in photochemical reaction

Rate constant (k) : It is equal to the rate of reaction when molar concentration of reactant is at unity.

Rate law : The rate of reaction is directly proportional to the product of concentration of reactant and each concentration is raised to the equal to no of moles actually participating in the reaction.

For a reaction, $aA + bB \rightarrow cC + dD$

$$\text{Rate law} = k[A]^p[B]^q$$

where powers p and q are determined experimentally.

Molecularity : The total number of reactants taking part in elementary chemical reaction is called molecularity.

Order of reaction : The sum of powers to which the concentration terms are raised in a rate law expression is called order of reaction.

For above case, $\text{Order} = P + Q$

Orders of reaction is determined experimentally.

Half-life period : The time during which the concentration of the reactant is reduced to half of its initial concentration is called half-life period.

Activation energy : The minimum extra amount of energy absorbed by reactant molecules so that their energy becomes equal to the threshold energy is called activation energy.

$$\text{Activation energy} = \text{Threshold energy} - \text{Kinetic energy}$$

Temperature coefficient : The ratio of rate constant at two temperatures having difference of 10°C is called temperature coefficient.

$$\text{Temperature coefficient} = \frac{\text{Rate constant at } T + 10^\circ\text{C}}{\text{Rate constant at } T^\circ\text{C}}$$

Arrhenius Equation :

$$K = Ae^{-E_a/RT}$$

where, K = Rate constant

A = Arrhenius energy (Frequency factor or pre-exponential factor)

E_a = Activation energy

R = Rate constant

T = Temperature

$$\frac{E_a}{RT} = \text{Fraction of molecules having energy equal to or more than activation energy}$$

$$\log K = \log A - \frac{E_a}{2.303RT}$$

$$\log \frac{K_2}{K_1} = \frac{E_a}{2.303R} \left[\frac{1}{T_1} - \frac{1}{T_2} \right]$$

$$\log \frac{K_2}{K_1} = \frac{E_a}{2.303 \times 8.314} \left[\frac{T_2 - T_1}{T_2 - T_1} \right] \quad \text{Where } R = 8.314 \text{ J mol}^{-1} \text{ K}^{-1}$$

$$\log \frac{K_2}{K_1} = \frac{E_a}{19.147} \left[\frac{T_2 - T_1}{T_2 - T_1} \right] = 0.0522 E_a \left[\frac{T_2 - T_1}{T_2 - T_1} \right]$$

1. Integrated rate law equation for zero order reaction is given as below :

$$(a) \quad k = \frac{[R]_0 - [R]_t}{t}$$

Where k is rate constant and $[R]_0$ is initial molar concentration.

$$(b) \quad t_{1/2} = \frac{[R]_0}{2k}, \quad t_{1/2} \text{ is half-life period of zero order reaction.}$$

2. Integrated rate law equation for first order reaction :

$$(a) \quad k = \frac{2.303}{t} \log \frac{[R]_0}{[R]_t}$$

where k is rate constant, $[R]_0$ is initial molar concentration and $[R]_t$ is final concentration at time ' t '.

$$(b) \quad \text{Half-life period } (t_{1/2}) \text{ for first order reaction :}$$

$$t_{1/2} = \frac{0.693}{k}$$

Pseudo chemical reaction : The chemical reaction which looks like higher order reaction but in real it follows lower order reaction.



$$\text{Rate} = k[\text{CH}_3\text{COOC}_2\text{H}_5]^1$$

$$\text{Order} = 1$$

VERY SHORT ANSWER TYPE QUESTIONS (1 Mark)

Q. 1. The gas phase decomposition of acetaldehyde



follows the rate law, $\text{Rate} = k(\text{P}_{\text{CH}_3\text{CHO}})^{3/2}$, unit of P is atm and time is second.

What are the units of its rate constant ?

Ans. $\text{Atm}^{-1/2} \text{ sec}^{-1}$

Q. 2. State the order with respect to each reactant and overall reaction.

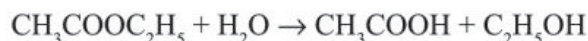


$$\text{Rate} = k[\text{H}_2\text{O}_2]^1[\text{I}^-]^1$$

Ans. Order of reaction = $1 + 1 = 2$

Q. 3. Give one example of pseudo first order reaction.

Ans. Hydrolysis of an ester :



Q. 4. The rate law for a reaction is

$$\text{Rate} = K [\text{A}] [\text{B}]^{3/2}$$

Can the reaction be an elementary process ? Explain.

Ans. No, an elementary process would have a rate law with orders equal to its molecularities and therefore must be in integral form.

Q. 5. For the reaction $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$, how are the rate of reaction expressions

$$-\frac{d[\text{H}_2]}{dt} \text{ and } \frac{d[\text{NH}_3]}{dt} \text{ inter-related ?}$$

$$\text{Ans. } -\frac{1}{3} \frac{d[\text{H}_2]}{dt} = \frac{1}{2} \frac{d[\text{NH}_3]}{dt}$$

Q. 6. Identify the order of a reaction from the following rate constant :

$$k = 2.3 \times 10^{-5} \text{ L mol}^{-1} \text{ s}^{-1}$$

Ans. Second order

Q. 7. After five half-life periods for a first order reaction, what fraction of reactant remains ?

$$\text{Ans. } \frac{1}{32}$$

Q. 8. Give the damaging effect of photochemistry.

Ans. The colour of fabrics fades away on exposure to sunlight.

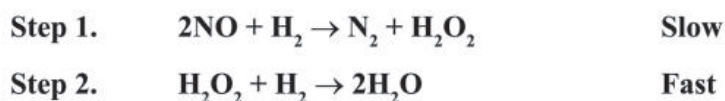
Q. 9. What is the effect of adding catalyst on the free energy of a reaction ?

Ans. No change in ΔG .

Q.10. What value of k is predicted for the rate constant by Arrhenius equation is $T \rightarrow \infty$? Is this value physically reasonable ?

Ans. From the equation $k = Ae^{-E_a/RT}$ if $T \rightarrow \infty$ $k \rightarrow A$ so that $E_a = 0$. This is not feasible.

Q.11. Determine the order of reaction :



Ans. $\text{Rate} = k [\text{NO}]^2 [\text{H}_2]$
 $\text{Order} = 2 + 1$
 $= 3$

Q.12. What is the order of reaction whose rate constant has the same units as the rate of reaction ?

Ans. Zero order

Q.13. Why are reactions of higher order less in number ?

Ans. A reaction takes place due to collide of molecules. The chances for a large number of molecules or ions to collide simultaneously are less. Hence, the reactions of higher order are less.

Q.14. What will be the effect of temperature on rate constant ?

Ans. Rate constant of a reaction is nearly doubled with rise in temperature by 10° .

Q.15. State a condition under which a bimolecular reaction is kinetically first order reaction.

Ans. A bimolecular reaction becomes first order reaction when one of the reactants is in excess.

Q.16. Why can't molecularity of any reaction be equal to zero ?

Ans. Molecularity of a reaction means the number of molecules of the reactants taking place in an elementary reaction. Since at least one molecule must be present, so that molecularity will be atleast one.

Q.17. Give an example of zero order reaction.

Ans. $\text{H}_2 + \text{Cl}_2 \xrightarrow{h\nu} 2\text{HCl}$ or any other example

Q.18. The rate constant of a reaction is $3 \times 10^2 \text{ min}^{-1}$. What is its order of reaction ? (On the basis of units of rate constant)

Ans. First order reaction.

Q.19. Three-fourth of a reaction is completed in 32 minutes. What is the half life period of this reaction ?

Ans. 16 minutes.

Q.20. What is meant by an elementary reaction ?

Ans. A reaction which takes place in one step is called an elementary reaction. For example : $\text{H}_2 + \text{I}_2 \rightarrow 2\text{HI}$.

Q.21. Give one example of a reaction where order and molecularity are equal ?

Ans. $2\text{HI} \rightarrow \text{H}_2 + \text{I}_2$ (Order = Molecularity = 2)

Q.22. For a reaction $R \rightarrow P$, the rate becomes 2 times when the concentration of the reactant A is increased 4 times. What is the order of reaction ?

Ans. $r = k(a)^n \Rightarrow 2r = k(4a)^n \Rightarrow 2 = 4^n \Rightarrow n = 0.5$

Q.23. The rate constant of a zero order reaction in A is $0.003 \text{ mol L}^{-1} \text{ sec}^{-1}$. How long will it take for the initial concentration of A to fall from 0.10M to 0.075 M ?

Ans. $t = \frac{[A]_0 - [A]}{k} = \frac{0.10 - 0.075}{0.003} = 8.3 \text{ sec}$

Q.24. In a reaction $2A \rightarrow \text{Products}$, the concentration of A decreases from 0.5 mol L^{-1} in 10 minutes. Calculate the rate during this interval.

Ans. Average rate $= \frac{-\Delta[A]}{2\Delta t} = -\frac{1}{2} \left(\frac{0.4 - 0.5}{10} \right) = 5 \times 10^{-3} \text{ M min}^{-1}$

Q.25. In some cases large number of colliding reactant molecules have energy more than threshold energy even then the reaction is slow. Why ?

Ans. Because resultant molecules do not collide in proper orientation.

Q.26. Give an example of a reaction having fractional order.

Ans. Decomposition of acetaldehyde (order = 1.5).



Q.27. Decomposition reaction of ammonia on Pt surface has rate constant $= 2.5 \times 10^{-1} \text{ mol L}^{-1} \text{ sec}^{-1}$. What is order of reaction ?

Ans. Unit of k explain that it is zero order reaction.

Q.28. What is order of radioactive decay ?

Ans. First order

Q.29. For a reaction $A + B \rightarrow \text{Product}$, the rate law is given by $r = k[A]^{1/2} [B]^2$. What is the order of the reaction ?

Ans. Order of reaction $= \frac{1}{2} + 2 = 2.5$

SHORT ANSWER-I TYPE QUESTIONS (2 Marks)

Q. 1. The rate of a particular reaction quadruples when the temperature changes from 293 K to 313 K. Calculate activation energy.

Ans. $K_2/K_1 = 4$

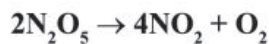
$$T_1 = 293 \text{ K}, T_2 = 313 \text{ K}$$

$$\log \frac{K_2}{K_1} = -\frac{E_a}{2.303R} \left[\frac{1}{T_1} - \frac{1}{T_2} \right]$$

Thus, on calculating and substituting values, we get :

$$E_a = 52.86 \text{ KJ mol}^{-1}$$

Q. 2. If the decomposition of nitrogen oxide as



follows a first order kinetics.

(a) Calculate the rate constant for a 0.05M solution if the instantaneous rate is $1.5 \times 10^{-6} \text{ mol/l/s}$?

Ans. $\text{Rate} = K[\text{N}_2\text{O}_5]$

$$K = \frac{\text{Rate}}{[\text{N}_2\text{O}_5]}$$

$$= \frac{1.5 \times 10^{-6}}{0.05}$$

$$K = 3.0 \times 10^{-5}$$

(b) What concentration of N_2O_5 would give a rate of $2.45 \times 10^{-5} \text{ mol L}^{-1} \text{ s}^{-1}$?

Ans. $\text{Rate} = 2.45 \times 10^{-5} \text{ mol L}^{-1} \text{ s}^{-1}$

$$[\text{N}_2\text{O}_5] = \frac{\text{Rate}}{K} = \frac{2.45 \times 10^{-5}}{3.0 \times 10^{-5}}$$

$$= 0.82 \text{ M}$$

Q. 3. Write the difference between order and molecularity of reaction.

Ans.

Order	Molecularity
1. It is the sum of the powers of concentration terms in the rate law expression.	It is the number of reacting species undergoing simultaneously collision in a reaction.
2. It is determined experimentally.	2. It is a theoretical concept.
3. Order of reaction need not to be a whole number.	3. It is whole number only.
4. Order of reaction can be zero.	4. It can't be zero or fractional.

Q. 4. The rate constant for a reaction of zero order A is $0.0030 \text{ mol L}^{-1} \text{ s}^{-1}$. How long will it take for the initial conc. of A to fall from 0.10M to 0.075M ?

Ans. For a zero order reaction,

$$K = 0.0030 \text{ mol L}^{-1} \text{ s}^{-1}$$

$$[A]_0 = 0.10\text{M} \quad [A] = 0.075\text{M}$$

$$K = \frac{[A]_0 - [A]}{t}$$

$$t = \frac{0.10 - 0.075}{0.0030} = 8.33 \text{ s}$$

Q. 5. For the reaction :



(a) Rate of reaction expression.

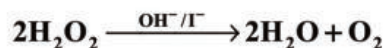
(b) Molecularity

Ans. (a) $\text{Rate} = -\frac{d[\text{C}_{12}\text{H}_{22}\text{O}_{11}]}{dt} = -\frac{d(\text{H}_2\text{O})}{dt}$

$$= \frac{d[\text{C}_6\text{H}_{12}\text{O}_6]}{dt} = \frac{d[\text{C}_6\text{H}_{12}\text{O}_6]}{dt}$$

(b) Molecularity = 2

Q. 6. Consider the decomposition reaction :



This reaction takes place in two steps as given below :



(a) Determine rate law expression.

(b) Determine the order of reaction.

Ans. (a) $\text{Rate} = K[\text{H}_2\text{O}_2][\text{I}^-]$ because second step is rate determining step.

(b) $\text{Order} = 1 + 1 = 2$

Q. 7. The decomposition of hydrocarbon follows the equation $K = (4.5 \times 10^{11} \text{ s}^{-1}) e^{-28000\text{K}/T}$. Calculate E_a .

Ans. $K = (4.5 \times 10^{11} \text{ s}^{-1}) e^{-28000\text{K}/T}$

Comparing the equation with Arrhenius equation,

$$K = Ae^{-E_a/RT}$$

$$-\frac{E_a}{R} = -28000 \text{ K}$$

$$\begin{aligned} E_a &= 28000 \times 8.314 \\ &= 232192 \text{ J mol}^{-1} \end{aligned}$$

Q. 8. A reaction is of second order with respect to a reactant. How is the rate of reaction affected if the conc. of the reactant is reduced to half. What is the unit of rate constant for such a reaction ?

Ans.

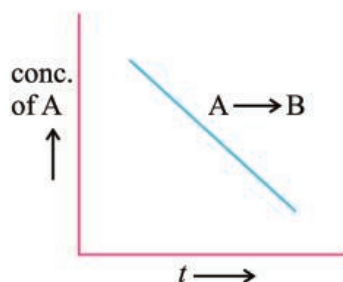
$$\text{Rate} = k[A]^2$$

$$\begin{aligned} \text{Unit of } k &= \frac{\text{mol L}^{-1}}{\text{S}} = k(\text{mol}^{-1})^2 \\ k &= \text{mol}^{-1} \text{ L S}^{-1} \end{aligned}$$

Q.9. For a first order reaction time taken for half of the reaction to complete is t_1 and $3/4$ of the reaction to complete is t_2 . How are t_1 and t_2 related ?

Ans. $t_2 = 2t_1$ because for $3/4$ th of the reaction to complete time required is equal to two half lives.

Q. 10.



(a) What is the order of the reaction ?

(b) What is the slope of the curve ?

Ans. (a) Zero order reaction.

$$(b) [R] = [R_0] - kt$$

$$\therefore \text{Slope} = -k$$

Q.11. Derive an expression to calculate time required for completion of zero order reaction.

Ans. For a zero order reaction,

$$R = [R]_0 - kt$$

For completion of the reaction $[R] = 0$

$$\therefore kt = [R]_0$$

$$\text{Or } t = \frac{[R]_0}{k}$$

Q.12. For the reaction $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

How is the rate of formation of ammonia related to the rate of disappearance of H_2 ?

$$\text{Ans. Rate of reaction} = -\frac{1}{3} \frac{\Delta[H_2]}{\Delta t} = \frac{1}{2} \frac{\Delta[NH_3]}{\Delta t}$$

$$\text{Or } \frac{\Delta[NH_3]}{\Delta t} = -\frac{2}{3} \frac{\Delta[H_2]}{\Delta t}$$

Q.13. The rate of a gaseous reaction becomes half when volume of the vessel is doubled. What is the order of reaction ?

Ans. Suppose, order of reaction is n and the reaction is $A(g) \rightarrow \text{Products}$

$$\text{Rate} = k[A]^n \quad \dots(i)$$

When volume is doubled, molar conc. becomes half and rate of reaction gets halved.

$$\frac{\text{Rate}}{2} = k \left(\frac{A}{2} \right)^n \quad \dots(ii)$$

Dividing equation (i) by equation (ii),

$$(2)^1 = (2)^n$$

$$\text{Or } n = 1$$

Q.14. A reaction which is first order with respect to A has rate constant 6 min^{-1} . If we start with $[A] = 0.5 \text{ mol L}^{-1}$, when would $[A]$ reach the value of 0.05 ML^{-1} ?

$$\text{Ans. } k = \frac{2.303}{t} \log \frac{[A]_0}{[A]}$$

$$k = 6 \text{ min}^{-1}, [A]_0 = 0.5, [A] = 0.05, t = ?$$

$$t = \frac{2.303}{6} \log \frac{0.5}{0.05} = \frac{2.303}{6} \log 10 = 0.3838 \text{ min}$$

Q.15. The conversion of the molecules X to Y follows second order kinetics. If the concentration of X is increased to three times, how will it affect the rate of formation of Y ?

Ans. 9 times

Q.16. A first order reaction has a rate constant $1.15 \times 10^{-3} \text{ s}^{-1}$. How long will 5 gram of this reactant take to reduce to 3 grams ?

Ans. $t = 444$ seconds

Q.17. $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$. If rate of formation of NO is $6 \times 10^{-4} \text{ atm min}^{-1}$, calculate the rate of formation of H_2O .

Ans. $9.0 \times 10^{-4} \text{ atm min}^{-1}$

Q.18. A first order reaction is 75% completed in 60 min. Find the half life of this reaction.

Ans. 30 mins.

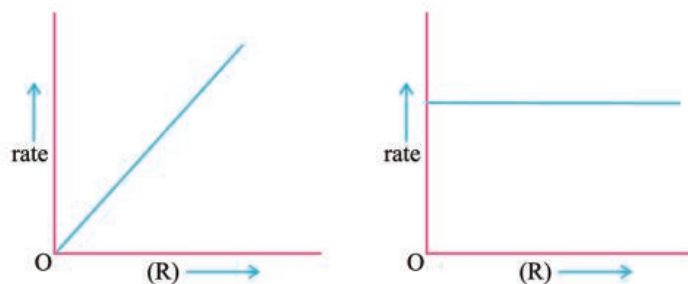
Q.19. The composition of a hydrocarbon follows the equation $K = (4.5 \times 10^{11} \text{ sec}^{-1}) e^{-28000\text{K}/T}$. Calculate the value of E_a .

Ans. 232.79 kJ/mol

Q.20. Consider a certain reaction $\text{A} \rightarrow \text{Product}$ with $K = 2.0 \times 10^{-2} \text{ s}^{-1}$. Calculate the concentration of A remaining after 100 s, if the initial concentration of A is 1.0 mol L^{-1} .

Ans. $[\text{A}] = 0.135 \text{ M}$

Q.21. Explain with an example, what is a pseudo first order reaction ? The graphs (A and B) given below are plots of rate of reaction Vs concentration of the reactant. Predict the order from the graphs.



Q.22. Differentiate between :

- (a) Average rate and instantaneous rate of a chemical reaction.
- (b) Molecularity and order of reaction.

Q.23. Show that in case of first order reaction, the time required for 99.9% of the reaction to take place is about ten times than that required for half the reaction.

Q.24. For the reaction $\text{NO}_2 + \text{CO} \rightarrow \text{CO}_2 + \text{NO}$, the experimentally determined rate expression below 400 K is $\text{rate} = k[\text{NO}_2]^2$. What mechanism can be proposed for this reaction ?

Q.25. The half life period of a first order reaction is 60 min. What % will be left after 240 mins. ?

Ans. 6.25%

Q.26. Time for half change for a first order reaction is 25 min. What time will be required for 99% reaction ?

Ans. 166.16 mins.

SHORT ANSWER TYPE-II QUESTIONS

Q. 1. The rate constant for first order reaction is 60/s. How much time will it take to reduce the concentration of the reaction to 1/10 of its initial value ?

Ans.

$$t = \frac{2.303}{K} \log \frac{[R_0]}{[R]}$$

$$t = \frac{2.303}{\frac{1}{10}} \log \frac{[R_0]}{[R]}$$

$$t = \frac{2.303}{60} \log 10$$

$$t = \frac{2.303}{60}$$

$$= 3.38 \times 10^{-2} \text{ s}^{-1}$$

Q. 2. The rate of most of reaction double when their temperature is raised from 298 K to 308 K. Calculate the activation energy of such a reaction.

Ans.

$$\log \frac{K_2}{K_1} = \frac{E_a}{2.303R} \left[\frac{1}{T_1} - \frac{1}{T_2} \right]$$

$$E_a = \frac{2.303 \times 8.314 \times 298 \times 308 \times 0.3010}{1000}$$

$$E_a = 52.89 \text{ KJ/mol}$$

Q. 3. A first order reaction takes 69.3 min for 50% completion. Set up an equation for determining the time needed for 80% completion.

Ans.
$$K = \frac{0.693}{t_{1/2}}$$

$$= \frac{0.693}{69.3} \text{ min}^{-1}$$

$$= 10^{-2} \text{ min}^{-1}$$

$$T = \frac{2.303}{K} \log \frac{[R_0]}{[R]}$$

$$= \frac{2.303}{10^{-2}} \log 5$$

$$= 160.9 \text{ min}$$

Q. 4. The activation energy of a reaction is 94.14 KJ/mol and the value of rate constant at 40° C is $1.8 \times 10^{-1} \text{ sec}^{-1}$. Calculate the frequency factor A.

Ans. Given, $E_a = 94.14 \times 10^3 \text{ J mol}^{-1}$, $T = 40 + 273 = 313 \text{ K}$, $K = 1.8 \times 10^{-1} \text{ sec}^{-1}$

By using, $K = Ae^{-E_a/RT} \Rightarrow \ln K = \ln A - \frac{E_a}{RT}$

Or $\log K = \log A - \frac{E_a}{2.303RT}$

Or $\log (1.8 \times 10^{-1}) + \frac{94.19 \times 10^3}{2.303 \times 8.314 \times 313} = \log A$

Or $A = \text{antilog}(10.9635) = 9.194 \times 10^{10} \text{ sec}^{-1}$

Q. 5. The rate constant of a reaction at 500 K and 700 K are 0.02 s^{-1} and 0.07 s^{-1} respectively. Calculate the value of E_a and A.

Ans. 18.23 KJ mol⁻¹, 1.603

Q. 6. The rate constant of a reaction at 700 K and 760 K are $0.011 \text{ M}^{-1} \text{ s}^{-1}$ and $0.105 \text{ M}^{-1} \text{ s}^{-1}$ respectively. Calculate the value of Arrhenius parameters.

Ans. 2.824×10^{10}

- Q. 7. The initial concentration of N_2O_5 in the first order reaction $\text{N}_2\text{O}_5 \rightarrow 2\text{NO}_2 + \frac{1}{2}\text{O}_2$ was $1.24 \times 10^{-2} \text{ mol L}^{-1}$ at 318 K. The concentration of N_2O_5 after 60 minutes was $0.20 \times 10^{-2} \text{ mol L}^{-1}$. Calculate the rate constant of the reaction at 318 K.

Ans.
$$K = \frac{2.303}{t} \log \frac{[A]_0}{[A]} = \frac{2.303}{t} \log \frac{[\text{N}_2\text{O}_5]_0}{[\text{N}_2\text{O}_5]_t} = \frac{2.303}{60} \log \frac{1.24 \times 10^{-2}}{0.2 \times 10^{-2}}$$

$$= \frac{2.303}{60} \log 6.2 = \frac{2.303}{60} \times 0.7924 \text{ min}^{-1}$$

$$= 0.0304 \text{ min}^{-1}$$

- Q. 8. The following data were obtained during the first order thermal decomposition of N_2O_5 at constant volume :



S. No.	Time per second	Total pressure (atm)
1	0	0.5
2	100	0.512

Calculate rate constant.

Ans. $4.98 \times 10^{-4} \text{ sec}^{-1}$

- Q. 9. A first order reaction is 20% complete in 20 minutes. Calculate the time taken for the reaction to go to 80% completion.

Ans. 144.3 minutes

- Q. 10. For a first order reaction, calculate the ratio between the time taken to complete $\frac{3}{4}$ of the reaction and the time taken to complete half of the reaction.

Ans. Two

- Q. 11. The following results have been obtained during the kinetics studies of the reaction :



Experiment	[A] mol L ⁻¹	[B] mol L ⁻¹	Initial rate of formation of D mol L ⁻¹ min ⁻¹
I	0.1	0.1	6.0×10^{-3}
II	0.3	0.2	7.2×10^{-2}
III	0.3	0.4	2.88×10^{-1}
IV	0.4	0.1	2.40×10^{-1}

Determine the rate law and the rate constant for the reaction.

Ans. Rate = $K[A][B]^2$

Q. 12. The rate of reaction triples when the temperature changes from 293 K to 313 K. Calculate the energy of activation of the reaction assuming that it does not change with temperature.

Q. 13. The decomposition of A into product has value of K as $4.5 \times 10^3 \text{ sec}^{-1}$ at 10°C and energy of activation 60 kJ/mol. At what temperature would K be $1.5 \times 10^4 \text{ sec}^{-1}$.

Ans. 24°C

Q. 14. (a) Write rate law and order of the following reaction :



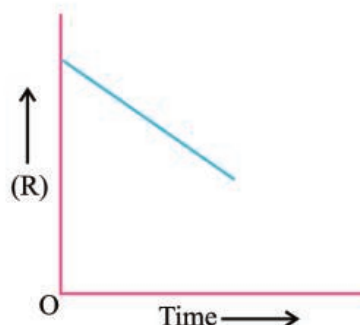
(b) Define energy of activation of a reaction.

(c) What is the relationship between rate constant and activation energy of a reaction ?

Ans. (a) Rate = $K[\text{AB}][\text{C}_2]$, Order = $1 + 1 = 2$

(c) $K = Ae^{-E_a/RT}$

Q. 15. For a chemical reaction $\text{R} \rightarrow \text{P}$, the variation in the concentration (R) vs time (t) plot is given :



(a) Predict the order of reaction.

(b) Write down its rate law.

(c) What is the slope of the curve ?

Q. 17. What do you understand by a first order reaction ? Show that for a first order reaction time required to complete a definite fraction of the reaction is independent of initial concentration.

Q. 18. In a pseudo first order reaction of hydrolysis of an ester in H_2O , the following results were obtained :

t/s	0	30	60	90
Ester (M/L)	0.55	0.31	0.17	0.085

- (a) Calculate the average rate of reaction between the time interval 30 to 60 sec.
- (b) Calculate the pseudo first order rate constant for the hydrolysis of ester.

Ans. (a) Average rate during 30-60 sec. $= \frac{0.17 - 0.31}{60 - 30} = 4.67 \times 10^{-3} \text{ mol L}^{-1} \text{ sec}^{-1}$

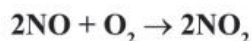
(b)
$$K_{30} = \frac{2.303}{t} \log \frac{[A]_0}{[A]} = \frac{2.303}{30} \log \frac{0.55}{0.31}$$

$$K_{60} = \frac{2.303}{60} \log \frac{0.55}{0.17}$$

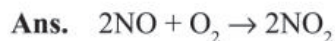
$$K_{90} = \frac{2.303}{90} \log \frac{0.55}{0.085}$$

$$\text{Average } K = 1.98 \times 10^{-2} \text{ sec}^{-1}$$

Q. 19. Following reaction takes place in one step :



How will the rate of the reaction of the above reaction change if the volume of reaction vessel is diminished to 1/3 of its original volume ? Will there be any change in the order of reaction with reduced volume ?



$$dx/dt = K[\text{NO}]^2[\text{O}_2]^1$$

[Since it is one step.]

If the volume of reaction vessel is diminished to 1/3, conc. of both NO and O₂ will become 3 times, the rate of reaction increased 27 times in the order of reaction with the reduced volume.

Q. 20. The decomposition of NH₃ on platinum surface is a zero order reaction. What are the rate of production of N₂ and H₂ ? [K = 2.5 × 10⁻⁴]



$$-\frac{1}{2} \frac{d[\text{NH}_3]}{dt} = \frac{d[\text{N}_2]}{dt} + \frac{1}{3} \frac{d[\text{H}_2]}{dt}$$

$$\begin{aligned} \frac{d[\text{NH}_3]}{dt} &= \text{rate} = k \times [\text{NH}_3]^0 \\ &= 2.5 \times 10^{-4} \text{ mol L}^{-1} \text{ sec}^{-1} \end{aligned}$$

$$\begin{aligned}\frac{d[\text{N}_2]}{dt} &= -\frac{1}{2} \frac{d[\text{NH}_3]}{dt} \\ &= \frac{1}{2} \times 2.5 \times 10^{-4} \text{ mol L}^{-1} \text{ sec}^{-1}\end{aligned}$$

$$\begin{aligned}d[\text{H}_2] &= -\frac{3}{2} \frac{d[\text{NH}_3]}{dt} = \frac{3}{2} \times 2.5 \times 10^{-4} \\ &= 3.75 \times 10^{-4} \text{ mol L}^{-1} \text{ sec}^{-1}\end{aligned}$$

$$\begin{aligned}\text{Rate} &= -\frac{d[\text{NH}_3]}{dt} = k[\text{NH}_3]^0 \\ &= 2.5 \times 10^{-4} \text{ mol L}^{-1} \text{ sec}^{-1}\end{aligned}$$

Rate of production of $\text{N}_2 = 2.5 \times 10^{-4} \text{ mol L}^{-1} \text{ sec}^{-1}$

LONG ANSWER TYPE QUESTIONS

Q. 1. (a) Define order of reaction.

(b) Rates of reaction double with every 10° rise in temperature. If this generalization holds for a reaction in the temperature ranges 298 K to 308 K, what would be the value of activation energy for their reaction ?
 $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$.

Ans. (a) Order of Reaction : It is the sum of powers to which the conc. terms are raised in rate law expression.

$$(b) \quad \log \frac{K_2}{K_1} = \frac{E_a}{2.303R} \left[\frac{1}{T_1} - \frac{1}{T_2} \right]$$

Here, $T_1 = 298 \text{ K}$, $T_2 = 308 \text{ K}$, $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$

$$\frac{K_2}{K_1} = 2$$

$$\begin{aligned}\log 2 &= \frac{E_a}{2.303 \times 8.314} \left[\frac{1}{298} - \frac{1}{308} \right] \\ 0.3010 &= \frac{E_a}{2.303 \times 8.314} \left[\frac{10}{298 \times 308} \right] \\ E_a &= \frac{0.3010 \times 2.303 \times 8.314 \times 298 \times 308}{10}\end{aligned}$$

$$= 52898 \text{ J mol}^{-1}$$

$$= 52.898 \text{ KJ mol}^{-1}$$

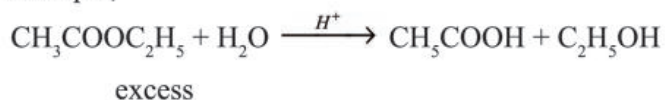
- Q. 2.** (a) What are pseudo order reaction ? Give example.
 (b) Rate constant K of a reaction varies with temperature 'T' according to the equation :

$$\log K = \log A - \frac{E_a}{2.303R} \left(\frac{1}{T} \right)$$

where E_a is the activation energy. When a graph is plotted for $\log K$ vs $1/T$, a straight line with a slope of -4250 K is obtained. Calculate E_a for the reaction.

- Ans.** (a) The chemical reaction which look like higher order reaction but in real they follow lower order kinetics.

For example,



$$\text{Rate} = K[\text{CH}_3\text{COOC}_2\text{H}_5]$$

$$\text{Order} = 1$$

$$(b) \quad \text{Slope} = \frac{E_a}{2.303R} = -4250 \text{ K}$$

$$\begin{aligned} \text{So, } E_a &= -2.303 \times R \times \text{Slope} \\ &= -2.303 \times 8.314 \text{ J K}^{-1} \text{ mol}^{-1} \times 4250 \\ &= 81375.3 \text{ J mol}^{-1} \\ &= 81.375 \text{ KJ mol}^{-1} \end{aligned}$$

- Q. 3.** (a) Determine the units of rate constant for first and zero order reaction.
 (b) Show that time required for the completion of 99% of the first order reaction is twice the 90% of completion of the reaction.

Ans. (a) $K = (\text{mol})^{1-n} \text{ L}^{n-1} \text{ S}^{-1}$

For zero order, $n = 0$

$$\text{So, } K = (\text{mol})^{1-0} \text{ L}^{0-1} \text{ S}^{-1} = \text{S}^{-1} \text{ mol L}^{-1}$$

For first order, $n = 1$

$$K = (\text{mol})^{1-n} \text{ L}^{n-1} \text{ S}^{-1}$$

$$\begin{aligned} \text{So, } K &= (\text{mol})^{1-1} \text{ L}^{1-1} \text{ S}^{-1} \\ &= \text{S}^{-1} \end{aligned}$$

(b) For a first order reaction,

$$t = \frac{2.303}{K} \log \frac{[A]_0}{[A]}$$

$$[A]_0 = a, [A] = a - \frac{a \times 99}{100} = 0.01 a$$

$$\begin{aligned} t(99\%) &= \frac{2.303}{K} \log \frac{a}{0.01a} \\ &= \frac{2.303}{K} \log 100 \\ &= \frac{2.303}{K} \times 2 \end{aligned} \quad \dots(i)$$

For 90% completion of reaction,

$$[A] = a - \frac{a \times 99}{100} = 0.1a$$

$$\begin{aligned} t(90\%) &= \frac{2.303}{K} \log \frac{a}{0.1a} \\ &= \frac{2.303}{K} \times 1 \end{aligned} \quad \dots(ii)$$

Dividing equation (i) by equation (ii), we get

$$t(99\%) = 2 \times t(90\%)$$

Q. 4. (a) Define rate constant of reaction.

(b) A first order reaction takes 40 mins for 30% decomposition. Calculate $t_{1/2}$.

Ans. (a) Rate constant : It is the rate of chemical reaction when the concentration of reactant taken as unity at a given temperature.

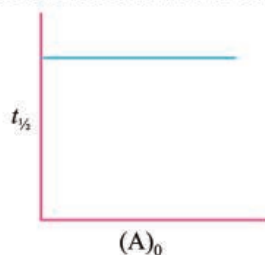
(b) Let initial conc. = a

$$\begin{aligned} \text{Conc. after 40 mins.} &= a - \frac{a \times 30}{100} \\ &= 0.70 a \end{aligned}$$

$$K = \frac{2.303}{t} \log \frac{[A]_0}{[A]}$$

$$\begin{aligned}
 &= \frac{2.303}{40} \log \frac{a}{0.70a} \\
 &= \frac{2.303}{40} \log \frac{1}{0.70} \\
 &= \frac{2.303}{40} \times 0.1549 \\
 &= 8.92 \times 10^{-3} \text{ min}^{-1} \\
 t_{1/2} &= \frac{0.693}{K} \\
 &= \frac{0.693}{8.92 \times 10^{-3}} = 77.7 \text{ min}
 \end{aligned}$$

- Q. 5. (a) Determine the order of reaction and also determine the units of rate constant.



- (b) The following data were given for thermal decomposition of SO_2Cl_2 at a constant volume :



Exp.	Time/s	Total p/atm
1	0	0.5
2	100	0.6

Calculate the rate of the reaction when total pressure is 0.65 atm.

- Ans. (a) First order reaction

$$K = \text{mol}^{1-n} \text{ L}^{n-1} \text{ S}^{-1}$$

$$n = 1$$

$$K = (\text{mol})^{1-1} \text{ L}^{1-1} \text{ S}^{-1}$$

$$= \text{S}^{-1}$$

- (b)

$$k = \frac{2.303}{t} \log \frac{P_i}{(2P_i - P_t)}$$

$$\begin{aligned}
 &= \frac{2.303}{100} \log \frac{0.5}{(2 \times 0.5 - 0.6)} \\
 &= \frac{2.303}{100} \log \frac{0.5}{0.4} \\
 &= \frac{2.303}{100} \times 0.969 = 2.23 \times 10^{-3} \text{ s}^{-1}
 \end{aligned}$$

Now, Rate = $P_{\text{SO}_2\text{Cl}_2}$

Pressure of SO_2Cl_2 when total pressure = 0.65 atm

$$\begin{aligned}
 P_{\text{SO}_2\text{Cl}_2} &= 2P_i - P_t \\
 &= 2 \times 0.5 - 0.65 \\
 &= 0.35 \text{ atm}
 \end{aligned}$$

$$\begin{aligned}
 \text{Rate} &= 2.23 \times 10^{-3} \times 0.35 \\
 &= 7.8 \times 10^{-4} \text{ atm S}^{-1}
 \end{aligned}$$

Q. 6. (a) The activation energy of a reaction is 100 kJ/mol. In the presence of catalyst the activation energy is decreased by 75%. What is the effect on rate constant of the reaction at 20° C ?

(b) $\text{A} + 2\text{B} \rightarrow 3\text{C} + 2\text{D}$

The rate of disappearance of B is $1 \times 10^{-2} \text{ mol L}^{-1} \text{ sec}^{-1}$. What will be (i) rate of reaction (ii) rate of change in the concentration of A and C ?

Ans. (a) 2.35×10^{13} times

$$\left[\text{Hint : } \log \frac{K_2}{K_1} = \frac{E_a - E_p}{2.303RT} \Rightarrow \log \frac{K_2}{K_1} = \frac{75 \times 10^3}{2.303 \times 8.314 \times 293} \right]$$

(b) $(5 \times 10^{-3}, 5 \times 10^{-3}, 15 \times 10^{-3}) \text{ M L}^{-1} \text{ sec}^{-1}$

Q. 7. (a) For the reaction $2\text{A} + \text{B} \rightarrow \text{A}_2\text{B}$

the rate = $k[\text{A}][\text{B}]^2$ with $k = 2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ sec}^{-1}$. Calculate the initial rate of reaction when $[\text{A}] = 0.1 \text{ mol L}^{-1}$, $[\text{B}] = 0.2 \text{ mol L}^{-1}$. Calculate the rate of reaction after $[\text{A}]$ is reduced to 0.06 mol L^{-1} .

(b) (i) The activation energy for the reaction $2\text{HI} \rightarrow \text{H}_2 + \text{I}_2$ is 209.5 KJ/mol at 581 K. Calculate the fraction of molecules of reactants having energy equal to or greater than activation energy.

(ii) The activation energy of a reaction is zero. Will the rate constant of the reaction depend upon temperature ? Give reason.

Ans. (b) (ii) $K = Ae^{-E_a/RT}$ if $E_a = 0$, then $K = A$. Frequency factor (A) does not depend upon temperature, therefore, rate constant and rate does not depend on temperature.

Q. 8. The concentration of R in the reaction $R \rightarrow P$ was measured as a function of time and the following data is obtained :

[R] molar	1.0	0.75	0.40	0.10
t (in min)	0.0	0.05	0.12	0.18

Calculate the order of the reaction.

[Hint : Determine the rate of reaction at different intervals of time. It comes out to be constant. Hence, it is a zero order reaction.]



